

Sadiq-ur-Rehman,<sup>a</sup> Saqib Ali<sup>b</sup>  
and Masood Parvez<sup>c\*</sup>

<sup>a</sup>Department of Chemistry, The University of Azad Jammu and Kashmir, Muzaffarabad, Pakistan, <sup>b</sup>Department of Chemistry, Quaid-i-Azam University, Islamabad-45320, Pakistan, and <sup>c</sup>Department of Chemistry, The University of Calgary, 2500 University Drive NW, Calgary, Alberta, Canada T2N 1N4

Correspondence e-mail: parvez@ucalgary.ca

## Key indicators

Single-crystal X-ray study  
 $T = 173$  K  
Mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å  
 $R$  factor = 0.042  
 $wR$  factor = 0.100  
Data-to-parameter ratio = 19.5

For details of how these key indicators were automatically derived from the article, see <http://journals.iucr.org/e>.

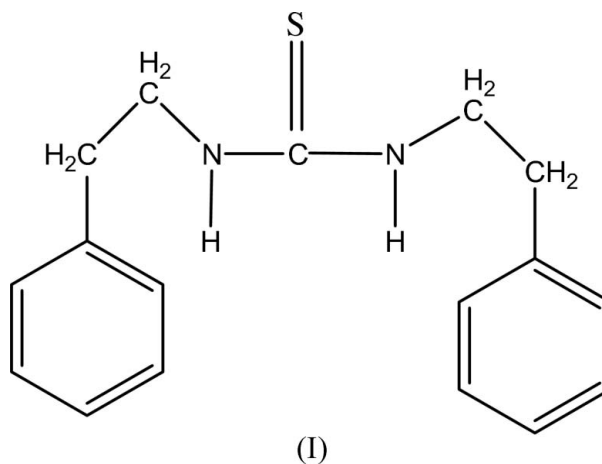
*N,N'*-Bis(2-phenylethyl)thiourea

The structure of the title compound,  $\text{C}_{17}\text{H}_{20}\text{N}_2\text{S}$ , exhibits hydrogen bonding of the type  $\text{N}-\text{H}\cdots\text{S}$ , resulting in chains of molecules running along the  $b$  axis.

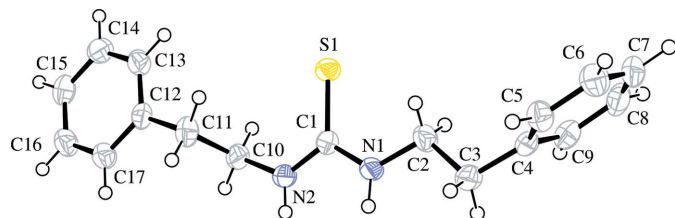
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## Comment

Thiourea derivatives are very useful building blocks for the synthesis of a wide range of aliphatic macromolecular and heterocyclic compounds. Benzothiazoles have been prepared from arylthioureas in the presence of bromine (Patil & Chedekel, 1984), while condensation of thiourea with halo-carbonyl compounds forms 2-aminothiazoles (Bailey *et al.*, 1996). *N*-substituted and *N,N'*-disubstituted thiourea derivatives have potential applications due to their coordination behaviour towards transition metals (Schuster *et al.*, 1990) and their biological activity (French *et al.*, 1970). For example, aliphatic and acylthioureas are well known for their fungicidal, antiviral, pesticidal and plant-growth regulating activities (Upadhyaya & Srivastava, 1982; Wegner *et al.*, 1986). Owing to the importance of thiourea, we report the synthesis and crystal structure of *N,N*-bis(2-phenylethyl)thiourea, (I).



The structure (Fig. 1) is composed of independent molecules of (I) forming chains along the  $b$  axis *via* hydrogen bonds involving amino H atoms and S1 atoms (Table 1 and Fig. 2). The molecular dimensions are as expected (CSD, Version 5.27; Allen, 2002). Atoms S1/N1/N2/C1/C3/C10 are almost coplanar, with a maximum deviation of 0.037 (1) Å for N2 from the plane formed by these atoms; atoms C2 and C11 are displaced by 0.145 (2) and 1.308 (2) Å, respectively, out of this plane. The C4–C9 phenyl ring is almost perpendicular to the above-mentioned plane, at an angle 85.43 (5)°. The C12–C17 phenyl ring is inclined at 9.45 (7)° with respect to the



**Figure 1**  
The molecular structure of (I), with displacement ellipsoids drawn at the 50% probability level.

mean plane of atoms S1/N1/N2/C1/C3/C10. The dihedral angle between the least-squares planes of the two phenyl rings is 76.09 (6)°.

## Experimental

A solution of 2-phenylethylamine (1.0 ml, 8.26 mmol) in acetone (20 ml) was added dropwise to a solution of CS<sub>2</sub> (0.5 ml, 8.26 mmol) and NH<sub>4</sub>OH (0.6 ml, 13 mmol) in acetone (20 ml). The mixture was stirred for about 4 h at room temperature. The solution was rotary evaporated under vacuum. The crude product was then poured into acidified water (200 ml) and stirred well. The solid product (I) was separated off and recrystallized from diethyl ether (yield 78%; m.p. 356 K).

### Crystal data

C <sub>17</sub> H <sub>20</sub> N <sub>2</sub> S	Z = 8
<i>M<sub>r</sub></i> = 284.41	<i>D<sub>x</sub></i> = 1.196 Mg m <sup>-3</sup>
Orthorhombic, <i>Pbca</i>	Mo <i>K</i> α radiation
<i>a</i> = 9.740 (2) Å	<i>μ</i> = 0.20 mm <sup>-1</sup>
<i>b</i> = 9.173 (2) Å	<i>T</i> = 173 (2) K
<i>c</i> = 35.354 (9) Å	Block, colourless
<i>V</i> = 3158.7 (12) Å <sup>3</sup>	0.20 × 0.16 × 0.14 mm

### Data collection

Nonius KappaCCD diffractometer	5741 measured reflections
<i>ω</i> and <i>φ</i> scans	3523 independent reflections
Absorption correction: multi-scan (SORTAV; Blessing, 1997)	2385 reflections with <i>I</i> > 2σ( <i>I</i> )
<i>T<sub>min</sub></i> = 0.962, <i>T<sub>max</sub></i> = 0.973	<i>R<sub>int</sub></i> = 0.037
	<i>θ<sub>max</sub></i> = 27.6°

### Refinement

Refinement on <i>F</i> <sup>2</sup>	$w = 1/[\sigma^2(F_o^2) + (0.029P)^2 + 0.81P]$
$R[F^2 > 2\sigma(F^2)] = 0.042$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.100$	( $\Delta/\sigma$ ) <sub>max</sub> = 0.001
<i>S</i> = 1.02	$\Delta\rho_{max} = 0.19 \text{ e } \text{Å}^{-3}$
3523 reflections	$\Delta\rho_{min} = -0.19 \text{ e } \text{Å}^{-3}$
181 parameters	
H-atom parameters constrained	

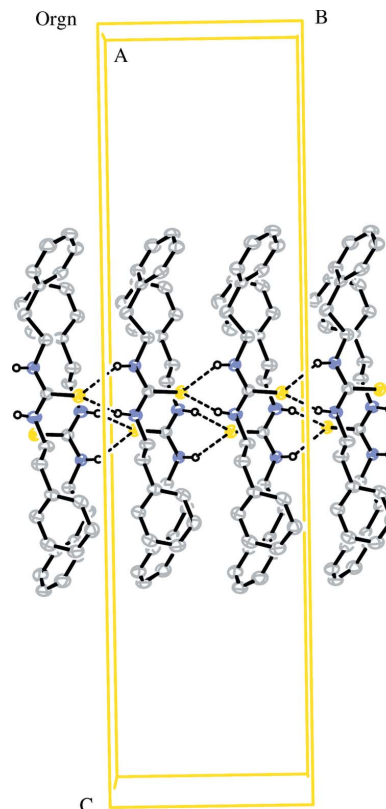
**Table 1**

Hydrogen-bond geometry (Å, °).

D—H...A	D—H	H...A	D...A	D—H...A
N1—H1...S1 <sup>i</sup>	0.88	2.81	3.563 (2)	144
N2—H2...S1 <sup>i</sup>	0.88	2.47	3.293 (2)	157

Symmetry code: (i)  $-x + \frac{3}{2}, y + \frac{1}{2}, z$ .

H-atoms were located in difference Fourier syntheses and were included in the refinement at geometrically idealized positions with N—H = 0.88, C—H = 0.95 and 0.99 Å and *U*<sub>iso</sub>(H) = 1.2*U*<sub>eq</sub>(C,N).



**Figure 2**  
Portion of the unit-cell contents of (I), showing hydrogen-bonded chains of molecules running along *b* axis. Dashed lines represent hydrogen-bonding interactions. H atoms not involved in hydrogen bonding have been omitted for clarity.

Data collection: *COLLECT* (Hooft, 1998); cell refinement: *DENZO* (Otwinowski & Minor, 1997); data reduction: *SCALE-PACK* (Otwinowski & Minor, 1997); program(s) used to solve structure: *SAPI91* (Fan, 1991); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ORTEPII* (Johnson, 1976); software used to prepare material for publication: *SHELXL97*.

## References

- Allen, F. H. (2002). *Acta Cryst.* **B58**, 380–388.  
 Bailey, N., Dean, A. W., Judd, D. B., Middlemiss, D., Storer, R. & Watson, S. P. (1996). *Bioorg. Med. Chem. Lett.* **6**, 1409–1414.  
 Blessing, R. H. (1997). *J. Appl. Cryst.* **30**, 421–426.  
 Fan, H.-F. (1991). *SAPI91*. Rigaku Corporation, Tokyo, Japan.  
 French, F. A., Blanz, E. J., DoAmaral, J. R. & French, D. A. (1970). *J. Med. Chem.* **13**, 1117–1124.  
 Hooft, R. (1998). *COLLECT*. Nonius BV, Delft, The Netherlands.  
 Johnson, C. K. (1976). *ORTEPII*. Report ORNL-5138. Oak Ridge National Laboratory, Tennessee, USA.  
 Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography, Part A*, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.  
 Patil, D. G. & Chedekel, M. R. (1984). *J. Org. Chem.* **49**, 997–1000.  
 Schuster, M., Kugler, B. & Konig, K. H. (1990). *Fresenius J. Anal. Chem.* **338**, 717–720.  
 Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.  
 Upadhyaya, J. S. & Srivastava, P. K. (1982). *J. Indian Chem. Soc.* **59**, 767–768.  
 Wegner, P., Kruger, H.-R., Franke, H. & Joppien, H. (1986). Eur. Patent No. 190 611.